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and a plot of $\ln(\)$ vs. should be linear for first order kinetics. The half-life can be given by $t_{1/2} = \frac{\ln(2)}{k}$ 2.2.3 N-TH ORDER KINETICS The above process can be done for any integer . We will consider the general reaction \rightarrow The rate law can be given by

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Octave Levenspiel, Chemical Reaction Engineering, 3rd Edition, Wiley, 1999. This is one of the gems of the chemical engineering literature. It was among the first textbooks of its kind back in 1962; ever since, its pedagogical value has been difficult to surpass.

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