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4

Chemistry-Study of change: Atomic
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and chemical reactions, molecular
structure, electronic configurations and
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elements, thermodynamics, states of matter, and properties of solutions. ...

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1. Define the following terms: (a) atom, (b) atomic number, (c) mass number, (d) isotopes (give 2 definitions), (e) atomic mass unit (amu), (f) nucleus, (g) spectroscopy, (h) ground state, (i) excited state, (j) atomic mass, (k) quantum leap, (l) photon.

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Chapter 4 Chemical Compounds - An
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These are the answers and explanations to the practice test on Chapters 1 - 3, which can be found here:
<https://goo.gl/NgVq75>

Chapter 4 An Introduction to Chemical Reactions

Chemistry Chapter 4 Study Guide 1. John Dalton's Atomic theory: - Element are made of tiny particles called atoms, All atoms of a given element are identical, The atoms of a given element are different from those of any other element , Atoms of one element can combine with atoms of other elements to form compounds.

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Terms in this set (31) "in any chemical or physical process, energy is neither created nor destroyed" is a statement of

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Democritus. Reasoned that atoms were indivisible and indestructible. Atom. The smallest particle of an element that retains its identity in a chemical reaction. Dalton. Transformed Democritus ideas on atoms into a scientific theory.

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1. Define the following terms: (a) atom, (b) atomic number, (c) mass number, (d) isotopes (give 2 definitions), (e) atomic mass unit (amu), (f) nucleus, (g) spectroscopy, (h) ground state, (i) excited state, (j) atomic mass, (k) quantum leap, (l) photon.

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answer the following questions. 27.

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Chapter 4 - Reactions in Aqueous
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description of how you will meet each
objective. (Although it is best to master
all of the objectives, the following

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has a net

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get a review of the most important
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selected problems at the end of the
chapter, and check your answers with
the solutions provided in this chapter of
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