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a. 2 and 4 c. 4 and 5 b. 3 and 4 d. 5 and 6 ____ 21. The discovery of the noble gases changed Mendeleev's periodic table by adding a new. a. period. c. group. b. series. d. sublevel block. ____ 22. In the modern periodic table, elements are ordered according to. a. decreasing atomic mass. c. increasing atomic number. b. Mendeleev's original ...

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HOMEWORK 5-1. 1. Explain how Mendeleev developed the periodic table. 2. Explain the term periodicity. How does the term apply to chemistry? Multiple Choice ____ 1. Which statement is true? a. The modern periodic table was developed at the First International Congress of Chemists. b. Motion needs to vary uniformly in order to vary periodically. c.

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The Periodic Law chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with the periodic law.

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Join a group of classmates from your inorganic chemistry course. Each group should have 3 or 4 students. Arrange a time to meet for an hour or so. Prior to the meeting, read part I (pp. 1367-1391) as well as the summary at the end of Linus Pauling's article (J. Am. Chem. Soc. 1931, 53, 1367-1400) on the nature of the chemical bond.

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