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A. 25.00 mL aliquot of concentrated hydrochloric acid (11.7 M) is added to 175.00 mL of 3.25 M hydrochloric acid. a. is this a dilution? b. determine the number of moles of hydrochloric acid from the 175.00 mL of 3.25 M hydrochloric acid. c. determine the number of moles of hydrochloric acid in the solution after the addition of the concentrated hydrochloric acid. assume the volumes are additive.

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### Redox Reactions Worksheet Instructional Fair

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3. What is the molarity of a solution of HNO<sub>3</sub> that contains 12.6 grams HNO<sub>3</sub> in 1.0 L of solution? ?  
mol HNO<sub>3</sub> = 12.6 g HNO<sub>3</sub> × 1 mol HNO<sub>3</sub> / 63.0 g HNO<sub>3</sub> = 0.200 mol HNO<sub>3</sub>  
M = 0.200 mol HNO<sub>3</sub> / 1.0 L = 0.200 M  
4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?  
0.700 M = moles of solute / 0.250 L  
moles of ...

#### **Molarity: Molarity = 1. 2.**

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What would be the molarity of this solution? The answer is 1.00 mol/L. Notice that both the units of mol and L remain. Neither cancels. A symbol for mol/L is often used. It is a capital M. So, writing 1.00 M for the answer is the correct way to do it. Some textbooks make the M using italics and some put in a dash, like this: 1.00-M.

#### **Molarity - ChemTeam**

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Molarity is expressed in units of moles per liter (mol/L). It's such a common unit, it has its own symbol, which is a capital letter M. A solution that has the concentration 5 mol/L would be called a 5 M solution or said to have a concentration value of 5 molar.