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Organic Chemistry Hydrocarbons Study Guide

Hydrocarbons An infinite variety of compounds can be assembled from only carbon and hydrogen atoms. Such hydrocarbons are the simplest organic compounds, but they are also of prime economic importance because they include the constituents of petroleum and natural gas.

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A SELF-STUDY GUIDE TO THE PRINCIPLES OF ORGANIC CHEMISTRY

Organic Chemistry Study Guide. STUDY. PLAY. unsaturated compounds. hydrocarbons containing carbon-carbon multiple

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bonds (double and triple bonds) ... (branched chain hydrocarbons) geometric isomers. each of two or more compounds that differ from each other in the arrangement of groups with respect to a double bond, ring, or other rigid ...

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Organic chemistry Today, the term organic compound is applied to all carbon-containing compounds with the primary exceptions of carbon oxides, carbides, and carbonates, which are considered inorganic. Because there are so many organic compounds, an entire branch of chemistry, called organic chemistry, is devoted to their study. Recall

Chapter 21: Hydrocarbons

Organic Chemistry Exam 1 Study Guide • “iso” structural unit: Carbon bonds to 1 Hydrogen and 2 CH₃ groups ex: CH₃CH₂CH₂CH₃ vs. CH₃CHCH₃ butane ï isobutane CH₃*isobutane

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is an isomer of butane because it contains an “iso” structural unit* • constitutional isomers consist of the same molecular formula but the atoms are connected differently than in the original molecule ...

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Chemistry – Hydrocarbons – Organic compounds are molecules comprising of both carbon and hydrogen atoms. These compounds can be found as gaseous, liquid, or solid. Organic compounds are vital in life processes as they are found in the products we use daily such as soap, cosmetics, perfumes, plastics, rubber, paper, insecticides.

Composition of Hydrocarbons ⇒ Free Book Summary

STUDY GUIDE In your textbook, read about organic chemistry, hydrocarbons, and straight-chain alkanes. Use each of the terms below just once to complete the passage. homologous series straight-chain alkanes If all of the carbon atoms are linked by single covalent bonds and there are no branches, the compounds are called (1) examples.

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Study Guide Key Concepts 22.1 Hydrocarbons • Because carbon has four valence electrons, carbon atoms always form four covalent bonds. • The carbon atoms in an alkane can be arranged in a straight chain or in a chain that has branches. • Molecules of hydrocarbons, such as alkanes, are nonpolar molecules. 22.2 Unsaturated Hydrocarbons

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