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Why are free radicals so reactive? Ask Question Asked 2 years, 11 months ago. Active 2 years, 11 months ago. Viewed 2k times 2. 2  $\begin{matrix} \text{ } \\ \text{ } \end{matrix}$  Why are free radicals are so reactive? They can break almost any bond, including  $\text{C-H}$  bonds, which are fairly stable. Don't they have an activation energy or something?

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## **organic chemistry - Why are free radicals so reactive ...**

Oxygen free radicals and their metabolites, collectively described as reactive oxygen species (ROS), have been implicated in the pathogenesis of many diseases. The pulmonary system is particularly vulnerable to ROS-induced

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injury because of its continuous exposure to toxic pollutants from a wide variety of sources in the ambient air.

### **The role of oxygen free radicals in occupational and ...**

The free radicals, both the reactive oxygen species (ROS) and reactive nitrogen species (RNS), are derived from

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both endogenous sources (mitochondria, peroxisomes, endoplasmic reticulum, phagocytic cells etc.) and exogenous sources (pollution, alcohol, tobacco smoke, heavy metals, transition metals, industrial solvents, pesticides, certain drugs like halothane, paracetamol, and radiation).

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What are free radicals? Well, in the body, free radicals typically take the form of reactive oxygen species, which include peroxides, superoxide, and the hydroxyl radical. Find more videos at

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## **Oxygen free radicals & cellular injury - causes, symptoms & pathology**

Highly reactive free radicals can damage important molecules like DNA, proteins, carbohydrates and lipids in the nucleus and in cell membranes. 6 Free radicals

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can also overwhelm the body and lead to inflammation, accelerated aging and oxidative aging. To combat free radicals, increase your intake of antioxidants from different sources (more on these to come later), since these may slow down ...

**Free Radicals Facts You Ought to**

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## **Know About**

In chemistry, a radical is an atom, molecule, or ion that has an unpaired valence electron. With some exceptions, these unpaired electrons make radicals highly chemically reactive. Many radicals spontaneously dimerize. Most organic radicals have short lifetimes.

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## **Radical (chemistry) - Wikipedia**

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The energetics of the benzoyl radical have been studied by photoacoustic

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calorimetry. The value obtained for D[PhC(O)-H],  $371 \pm 11 \text{ kJ mol}^{-1}$  ( $89 \pm 3 \text{ kcal mol}^{-1}$ ), agrees with an early result derived from iodination kinetic experiments and does not support a recent suggestion that the radical is resonance stabilized.

### **Enthalpy of formation of the**

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Cell-damaging free radicals. Free radicals are highly reactive molecules that travel around the body, attaching and binding to other molecules. This



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starts a destructive chain reaction, turning any molecule they come in contact with into an unstable free radical. ... Author J Hay. The Inhibition of Fat Oxidation Processes. Author N. M. Emanuel ...

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Free radicals are highly reactive and

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unstable molecules that are produced in the body naturally as a byproduct of metabolism (oxidation), or by exposure to toxins in the environment such as tobacco smoke and ultraviolet light. Free radicals have a lifespan of only a fraction of a second, but during that time can damage DNA, sometimes resulting ...

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## **Free Radicals: Definition, Cause, and Role in Cancer**

What Are Free Radicals? Free radicals are atoms, ions, or molecules that contain an unpaired electron. The unpaired electron makes them unstable and highly reactive. In a process called oxidation, free radicals steal electrons from other molecules — fats, proteins,

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cell membranes, and even DNA —  
altering the fundamental structure of the

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